FOCUS ACADEMY

Kg to 12
English&Gujarati Medium

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STD- 9TH SUB- SCIENCE CHAPTER 1&2

1} Gaseous state

- They do not have definite shape and take up the shape of the container.
- They do not possess definite volume due to weakest intermolecular forces.
- They are not rigid.
- They are easily compressible due to excess space between the particles of gas which compresses on applying pressure.
- They can easily undergo diffusion due to the fact that molecules in a gas moves at a very fast rate due to which speed of diffusion is very large.
- They can flow in all possible directions.

2} A phase is reached where at certain temperature solid gets converted to liquid. This temperature at which a solid melts and converts into liquid at atmospheric pressure is termed as melting point.

3} Particles of water possess greater kinetic energy at 0°C. On application of more heat the particles of water starts moving faster. At specific temperature the particles acquire enough energy to break free from the forces of attraction of each other. At this temperature the liquid starts changing into gas. This temperature at which a liquid starts boiling and starts converting to gas at the atmospheric pressure is called boiling point.

4} The Changing of a solid directly into vapours on heating, and of vapours into solid on cooling is called as sublimation. The solid substance which undergoes sublimation is called sublime.

5}The process of a liquid changing into vapour (or gas) even its boiling point is called evaporation.

6}A solution is **a homogeneous mixture of two or more substances**. ..

Colloids (also known as colloidal solutions or colloidal systems) are **mixtures in which microscopically dispersed insoluble particles of one substance are suspended in another substance**.

7}Chromatography is **an important biophysical technique that helps in separation, identification and purification of a compound from the given mixture**. .

8}Fractional distillation is **a process of separating two (or more) miscible liquids by distillation, the distillate being collected in fractions boiling at different temperatures**. ... For example: - Mixture of acetone and water can be separated by fractional distillation.

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